

ICE Problems

I Initial
C Change
E Equilibrium

WEAK ACIDS

$$K_a = \frac{[H^+(aq)]^2}{\text{conc of acid}}$$

[HA]

To find K_a you must use an Ice table and consider the concentration change that takes place because dissociation is not 100%

EXAMPLE: A 0.1000 M solution of ethanoic acid is only partially ionized. From a pH meter and calculation involving pH, the hydrogen ion concentration is found to be 1.34×10^{-3} M. Find the K_a using an ICE table.

